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Writing and naming compounds

Tyler DeWitt is back at it
again

<https://www.youtube.com/watch?v=URc75hoKGLY>

Some helpful Vocabulary

Binary Compounds – a chemical compound that consists of TWO elements.

Polyatomic Ion – an ion that contains more than one atom.

Binary ionic compound – a **compound** held together by ionic bonds.

Binary molecular compound – a **molecule** made up of covalently bonded atoms.

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Write the directions for solving binary compounds

Now lets practice (page 362 will help)

Ba and Cl

N and H

Al and O

You need to familiar with these

1. Ammonium NH_4^+
2. Carbonate CO_3^{2-}
3. Hydroxide OH^-
4. Nitrate NO_3^-
5. Peroxide O_2^{2-}
6. Phosphate PO_4^{3-}
7. Sulfate SO_4^{2-}

Oxidation Numbers

1+	2+	Most common oxidation number										3+	4+	3-	2-	1-	
Li 3	Be 4											B 5	C 6	N 7	O 8	F 9	He 2
Na 11	Mg 12											Al 13	Si 14	P 15	S 16	Cl 17	Ne 10
K 19	Ca 20	Sc 21	Ti 22	V 23	Cr 24	Mn 25	Fe 26	Co 27	Ni 28	Cu 29	Zn 30	Ga 31	Ge 32	As 33	Se 34	Br 35	Kr 36
Rb 37	Sr 38	Y 39	Zr 40	Nb 41	Mo 42	Tc 43	Ru 44	Rh 45	Pd 46	Ag 47	Cd 48	In 49	Sn 50	Sb 51	Te 52	I 53	Xe 54

NOTE: Many elements have more than one possible oxidation number.

NAMING *ionic* COMPOUNDS

1. Write the name of the first element
2. Write the root name of the second element
3. Add the suffix -ide to the root name

N and Cl

Mg and O

NAMING COMPOUNDS with *polyatomic ions*

1. Write the name of the first element of polyatomic ion first. Use the periodic table or ion chart on page 366 to find its name.
2. Write the name of the second element or polyatomic ion second. Use the periodic table or ion chart on page 366 to find its name. IF THE SECOND ONE IS AN ELEMENT, USE THE ROOT NAME OF THE ELEMENT AND ADD THE SUFFIX -IDE

Let's practice

1. Potassium and hydrogen and Sulfate
2. Manganese Carbonate
3. CaCO_3